

Naming Formulas-General Rules

Determine if the compound is ionic, covalent, or an acid. Then name!

Naming Ionic Compounds: metal (+) & nonmetal (-) bond or formula units

a. Cation (+)

1. Name of metal (or positive group: NH_4^{+1})
2. If two or more charges use a Roman No. (also know as ic and ous for the old naming system)

b. Anion (-)

1. Nonmetal -one type = _____ *ide*
2. Group-name the polyatomic ion group

Naming Covalent Compounds: nonmetal-nonmetal sharing or molecule

a. Prefix nonmetal - Prefix nonmetal-ide (if two different elements)

- | | |
|----------|----------|
| 1. mono | 6. hexa |
| 2. di | 7. hepta |
| 3. tri | 8. octa |
| 4. tetra | 9. nona |
| 5. penta | 10. deca |

Naming Acids: General form is H X (hydrogen is H^{+1}) (keep ur for sulfur and phosphorus ions only)

- a. If X = element then hydro_____ ic acid
- b. If X = group-ate then _____(ur)ic acid
- c. If X = group-ite the _____(ur)ous acid